to be active. However, the potency was not very high. In three cats the geometric mean lethal dose was found to be 1.806 mg./Kg.

Acospectoside A and acobioside A were also tested on isolated perfused guinea pig heart. Both the glycosides caused a positive ionotropic action. Results of preliminary testing suggested that in case of acospectoside A, increase in contractility was rapid in onset but brief in duration, and was approximately proportional to the dose. In a guinea pig heart, doses of 2 and 25 mcg. caused 110 and 167% increase of contraction, respectively.

Effect of acobioside A was comparatively slow in onset but persistent in duration. In the same heart it appeared to be 3–10 times more potent than acospectoside A.

(1) Kapadia, G. J., Kapadia, G. G., and Mosby, J. R., *Lloydia*, **27**, 233(1964).
(2) *Ibid.*, **27**, 272(1964).
(3) Euw, J. V., *et al.*, *Helv. Chim. Acta*, **34**, 1821(1951).
(4) Hauschild-Rogat, P., *et al.*, *ibid.*, **45**, 2116(1962).

GOVIND J. KAPADIA

Department of Pharmacognosy and Natural Products
College of Pharmacy
Howard University
Washington, D. C.

Received September 3, 1965.
Accepted for publication October 6, 1965.

This investigation was supported by grant HE-06571 from the National Institutes of Health, U. S. Public Health Service, Bethesda, Md.

This report is part of a paper presented to the American Society of Pharmacognosy, Pittsburgh meeting, June 1964. (See *Reference 2* for the published abstract.)

The author appreciates the kind cooperation of Professor T. Reichstein, University of Basel, and thanks him for furnishing samples of acovenosides A and B and acobioside A. The author also thanks Professor K. K. Chen, Indiana University, for toxicity assay of acospectoside A on cats; Professor S. N. Pradhan, Howard University, for tests on isolated guinea pig heart; and Dr. R. J. Highet and Mrs. K. S. Warren, National Heart Institute, for spectral data and optical rotation determinations.

Errata

In the article titled "Solubilization of Weakly Acidic and Basic Drugs by Aqueous Solutions of Polysorbate 80" (1), the sentence beginning on line 6, column 2, page 1348, should read:

The greater apparent energy released by partitioning of the ionized acid may be due, in part, to entropy contributions related to the ion–dipole interaction with water which does not occur in the case of the free acid.

(1) Rippie, E. G., Lamb, D. J., and Romig, P. W., *J. Pharm. Sci.*, **53**, 1346(1964).

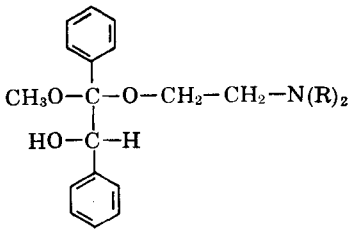
In the article titled "Biopharmaceutical Investigation of Nalidixic Acid in Man" (1), the following corrections should be made:

1. Under the section *In Vitro Dissolution Rates* on page 38, sentence 4 should read: The method consisted of using a phosphate buffer at pH 7.5 in a volume of 400 ml./500 mg. of nalidixic acid at 37°.
2. In Table III on page 40 under "... k_E Min.$^{-1}$ (Range)" for "Micropulverized powder—Caplet" change 0.00019 to 0.0019.
3. In Table IV on page 40 under "Min. Peak Time" for "Sodium salt capsule" change 46 to 48.

(1) Moore, W. E., *et al.*, *J. Pharm. Sci.*, **54**, 36(1965).

In the article titled "Synthesis and Antitremorine Activity of Amino Ketals" (1), the following corrections should be made:

1. On page 60, structure VI should be:



VI

2. On page 61, in Table II under "Compd.," change VI^a to VI*a*.
3. Page 63, column 2, line 3 should read: **1,2-Diphenyl-2-methoxy-2-(β-piperidinoethoxy)-ethanol Oxalate (VIc).—**

(1) Johnson, H. L., and Oneto, J. F., *J. Pharm. Sci.*, **54**, 59(1965).

In the article titled "Polymeric Pharmaceutical Coating Materials I. Preparation and Properties" (1), the following correction should be made on page 178:

The spectra in Figs. 3 and 4 should be transposed. The captions should remain the same.

(1) Lappas, L. C., and McKeehan, W., *J. Pharm. Sci.*, **54**, 176(1965).

In the article titled "Phytochemical Investigation of *Acacia angustissima*" (1), the following correction should be made:

In Table III on page 238, the color codes designated solely as F should be changed to P.

(1) Hammer, R. H., and Cole, J. R., *J. Pharm. Sci.*, **54**, 235(1965).

In the article titled "Solubility Profiles for the Xanthines in Dioxane–Water Mixtures" (1), the following corrections should be made:

1. The numbers should be added to the axes of Figs. 1–3 on page 839.
2. Figures 2 and 3 should be transposed.
3. The corrected figures are as follows:

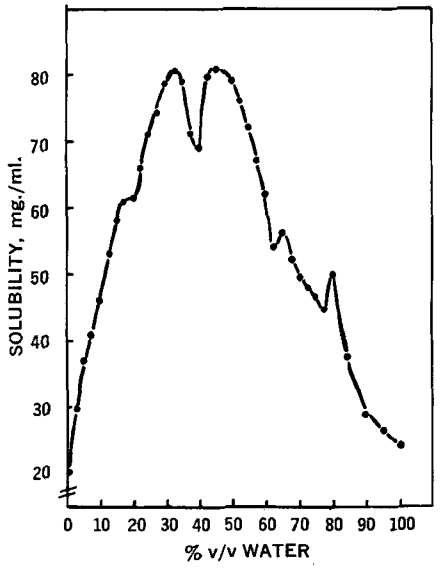


Fig. 1.—The solubility of caffeine at 25°C. in mg./ml. is plotted as a function of composition (v/v) for dioxane–water mixtures.

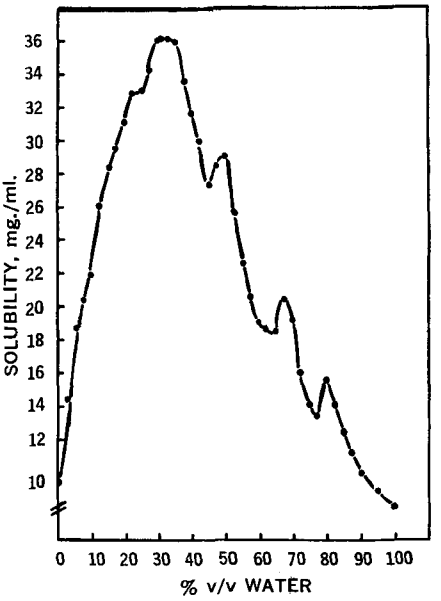


Fig. 2.—The solubility of theophylline at 25°C. in mg./ml. is plotted as a function of composition (v/v) for dioxane–water mixtures.

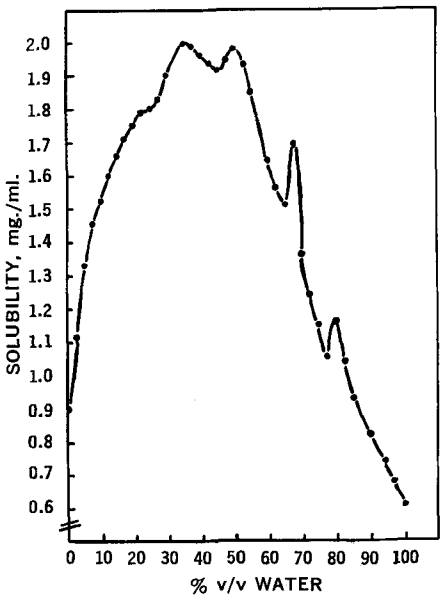


Fig. 3.—The solubility of theobromine at 25°C. in mg./ml. is plotted as a function of composition (v/v) for dioxane–water mixtures.

(1) Paruta, A. N., Sciarrone, B. J., and Lordi, N. G., *J. Pharm. Sci.*, **54**, 838(1965).